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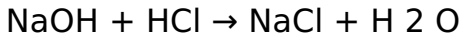
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Young Wonders ~~Acids Bases And~~
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Solution: The reaction of the acid + base gives a product of salt + water, which is considered as neutralization reaction. Examples:



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Short Answer Type. Question 1.
Two solutions A and B have pH
values of 5 and 8 respectively.
Which solution will be basic in
nature ...

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Ans:- Salts are classified into four categories on the basis of the strength of acid and base. they are: Salt of strong acid and strong base.e.g NaCl, KCl; Salt of strong acid and weak base. e.g. NH_4Cl ,

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NH_4I ; Salt of weak acid and strong base. e.g. CH_3COONa , $\text{COO}^- \text{Na}^+$ + Salt of weak acid and weak base. e.g. $\text{CH}_3\text{COO}^- \text{NH}_4^+$ + How is common salt found in nature?

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2. Bases react with acids to produce salts and water; water to produce acids and salts; salts to produce acids and water; neither acids, salts, nor water

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The salts of strong acids and weak bases give acidic solution having pH less than 7. Example, NH_4Cl , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7. Example, Na_2CO_3 ,

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Sodium Carbonate will have pH
more than 7. Question 7.

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Answer A reaction in which an
acid and base react with each

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other to give a salt and water is termed as neutralization reaction. For Example: (i) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ (ii) $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$ 15. Give two important uses of washing soda and baking soda. Answer Two important uses of washing

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Answer. (a) HCl is acid and NaOH
is base whose combination forms
the common salt. Its formula is
NaCl (Sodium chloride). It is

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Chemistry
obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl . It may be white or light blue or yellow depending upon impurities present in it.

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(A) inorganic acid (B) organic acid
(C) soft organic acid (D) strong
inorganic acid. Answer. Answer:

(a) 5. When milk of magnesia
reacts with acetic acid it produces
(A) basic salts (B) acidic salts (C)
neutral salt (D) complex salt.

Answer. Answer: (c) 6. Which of

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the following phenomena will
occur when a small amount of
acid is added to water ...

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Turmeric is a natural indicator. On adding its paste to acid and base separately, which colours would be observed. (a) Yellow in both acid and base (b) Yellow in acid and red in base (c) Pink in acid

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and yellow in base (d) Red in acid
and blue in base. Answer.

Answer: (b) Yellow in acid and red
in base

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Question 1.

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Chemistry
a look at one vital part of the subject - acids, alkalis, bases and salts. In it we find out the difference between alkalis and bases, how acids and alkalis react with one another, how neutralisation can produce salts and the relevance of pH values.

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Acids, bases and salts An acid is a compound that is defined by its physical and chemical properties. Acids taste sour and react with metals and polyatomic ions called

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Carbonates. A carbonate is a charged cluster of Carbon and Oxygen atoms.

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Answers to help students
understand the concept very well.

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Acid + Base \rightarrow Salt + Water.

Question 3: Rain becomes acidic
due to. carbon dioxide. sulphur

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Chemistry
dioxide, nitrogen dioxide. all of
the above. Correct Option is : 4.
Solution : Rain becomes acidic
because of carbon dioxide,
sulphur dioxide and nitrogen
dioxide. They dissolve in rain
drops to form carbonic acid,
sulphuric acid and nitric acid ...

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