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ANSWER Answer the

following questions

in the space

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provided. 1. Given the following equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? 40.07 g/mol

b. What is the molar mass of C_3H_8 ? 2 mol O_2 :1 mol H_2O

c. What is the mole ratio of O_2 to H

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molecules of

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hydrogen peroxide
(in aqueous solution)
decompose to
produce two
molecules of liquid
water and one
molecule of oxygen
gas. Chapter 9:
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SECTION 3 PROBLEMS

Write the answer on
the line to the left.

Show all your work in

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the space provided.

1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.
2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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Theoretically, how many moles of NH_3 will be produced?

PROBLEMS Write the answer on the line to the left, Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are

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mixed with 12.0 mol
of H₂ according to
the ...

Answers

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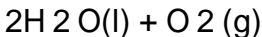
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example is included
below: $2\text{H}_2\text{O}_2(\text{aq})$



This describes the
decomposition
reaction of hydrogen
peroxide.

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molecules of
hydrogen peroxide
(in aqueous solution)
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with mass

relationships of

elements in

compounds Reaction

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Involves mass

relationships

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a. What is the value of the coefficient x in this equation? 40.07

g/mol b. What is the molar

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ANSWER Answer the following questions in the space

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