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Energy, Transition

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890.4 kJ/mol of heat.

That is, when one mole of methane,  $\text{CH}_4$ , is burned, 890.4 kJ of heat are given off to the surroundings.

This means that the products have 890.4 kJ less energy stored in the bonds than the reactants.

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Thus,  $\Delta H$  for the  
reaction =  $-890.4$  kJ.

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890.4 kJ/mol of heat.

That is, when one 1. The  
combustion of methane,

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CH<sub>4</sub>, releases mole of methane is burned, 890.4 kJ are given off to the surroundings This means that the products have 890.4 kJ less energy stored in the bonds than the reactants.

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enthalpy of atomization

( $\Delta_{\text{atom}}H$ ):  $\Delta_{\text{atom}}H =$

$H(\text{products}) -$

$H(\text{reactants}) = H(\text{atoms})$

$- H(\text{compound})$  (1)

enthalpy,  $H$   $H_{\text{products}}$

reactants  $H_{\text{reactant}}$

products  $\Delta_{\text{atom}}H$  ( $H_2$ ) =

+435 kJ 1 mole of  $H_2(g)$

2 moles of  $H(g)$  When a

mole of a compound is

made from its

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constituent gas phase  
atoms, energy is  
released and the energy  
change is called the  
enthalpy of

**Enthalpy Of Reaction (**  
**H CHEM1101**  
**Worksheet 10:**  
**Enthalpy ...**

In this worksheet,  
students will look at  
enthalpy graphs and  
measure the activation

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energy of a reaction.

They will also see the effect that a catalyst has on activation energy.

Essential

concepts: Energy, heat, enthalpy, activation energy, potential energy, exothermic, endothermic.

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enthalpy from enthalpy change of formation.

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Answer the following to the best of your ability.

Questions left blank are not counted against you.

When you have completed every question that you desire, click the " MARK TEST " button after the last exercise.

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on the state function  
status of enthalpy. You  
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values of known reactions to find the  $\Delta H$  of a new reaction. How do the following changes to a reaction affect  $\Delta H$  values?

Reverse of reaction  
 $\Delta H$  Reverses sign of  $\Delta H$

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Enthalpy depends upon the amount of

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substance; i.e., it is an extensive property. Since thermochemical equations are read with the understanding that the coefficients are numbers of moles, changing the coefficients by multiplying through the equation also multiplies the value of the enthalpy.



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Answers 6 h. I i. I j. C k.

N l. C m. C n. N o. IN p.

C q. I 2. For the

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standard molar enthalpy  
of formation (of the  
solid compound), the  
enthalpy of sublimation  
of the metal, the

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