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Calculations involving
Specific Heat.

Worksheet-

Calculations involving
Specific Heat. 1. For
 $q = m c \Delta T$: identify
each variables by
name & the units
associated with it. $q =$

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amount of heat (J) m

= mass (grams) c =

specific heat ($\text{J/g}^\circ\text{C}$)

ΔT = change in

temperature ($^\circ\text{C}$) 2.

Heat is not the same
as temperature, yet
they are related.

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aluminum skillet
weighing 1.58 kg is

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heated on a stove to 173 °C. Suppose the skillet is cooled to room temperature, 23.9 °C. How much heat energy (joules) must be removed to cause this cooling? The specific heat of aluminum is 0.901 J/(g · °C). 2.

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Solving For Heat (q)

1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 °C to 18.0 °C?
2. How much heat is lost when a 64 g piece of copper cools from 375 °C, to 26 °C? (The specific heat of copper is 0.38452 J/g

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x °C). Place your answer in kJ. 3. The specific heat of iron is 0.4494 J/g x °C. How much heat is transferred when a 4.7 kg piece

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Heat. DIRECTIONS:

Use $q = (m)(\Delta T)(C_p)$ to solve the following problems. Show all

work and units. A

15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

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Specific Heat Worksheet

Two page worksheet using Specific Heat Capacity. Questions start easy then become gradually harder. Answers included on separate sheet. Also includes a spreadsheet to show how the calculations have been done.

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Specific Heat
Capacity Worksheet
(with answers) |

Teaching ...

Latent heat and
Specific heat capacity
questions. 1. How
much water at 50°C
is needed to just melt
2.2 kg of ice at 0°C ?
2. How much water at
 32°C is needed to

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just melt 1.5 kg of ice

at -10°C ? 3. How

much steam at 100°

is needed to just melt

5 kg of ice at -15°C ?

4. A copper cup holds

some cold water at

4°C .

Latent heat and

Specific heat capacity

questions.

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HEAT Practice
Problems . $Q = m \times$
 $c \times \Delta T$. 5.0 g of
copper was heated
from 20°C to 80°C.
How much energy
was used to heat Cu?
(Specific heat
capacity of Cu is

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Specific Heat

0.092 cal/g °C) 27.6 cal. How much heat is absorbed by 20g granite boulder as energy from the sun causes its temperature to change from 10°C to 29°C? (Specific heat capacity of granite is 0.1 cal/g°C) 38 cal

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Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units. Ex:
How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if

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Worksheet
Before discussing
Calculating Specific
Heat Worksheet
Answers, you need to
recognize that
Knowledge can be

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your answer to a better the next day, along with studying doesn't just stop the moment the school bell rings. Of which getting claimed, many of us provide you with a a number of basic yet helpful posts along with design templates made ideal for almost any educative purpose.

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CH 8: Specific Heat
Problems Worksheet.

1. How much energy
must be absorbed by
20.0 g of water to
increase its
temperature from
283.0 °C to 303.0

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°C? 2. When 15.0 g of steam drops in temperature from 275.0 °C to 250.0 °C, how much heat energy is released? 3.

Thermochemistry
Problems - Worksheet
Number One
Calorimetry Practice
Problems 1. How
much energy is

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needed to change the temperature of 50.0 g of water by 15.0°C?

2. How many grams of water can be

heated from 20.0 °C to 75°C using

12500.0 Joules? 3.

What is the final

temperature after 840 Joules is absorbed by

10.0g of water at

25.0°C? 4. The heat capacity of aluminum

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is $0.900 \text{ J/g}^\circ\text{C}$. a.

Problems

Worksheet With Answers Calorimetry Practice Problems

If the specific heat of water is $4.18 \text{ J/g}^\circ\text{C}$, calculate the amount of heat energy needed to cause this rise in temperature.

Specific Heat (C):
 0.03 A total of 54.0
Joules of heat are

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the end of this
worksheet to solve
this problem. Is
energy absorbed or
released? ... The
specific heat of liquid

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ethanol is $2.44 \text{ J/g}^\circ\text{C}$.

6. How much energy in joules does 28.5g of liquid sulfur lose when it lowers from 120°C to 115°C , then change into a solid?

The specific heat of liquid sulfur is $0.71 \text{ J/g}^\circ\text{C}$ More

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understand how to

project cash flow.

Whatever your

company planning

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objectives, cash flow is still the resource in the organization, and managing money is the business purpose. Version control is another significant issue with Excel.

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templates they
provide or Start
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